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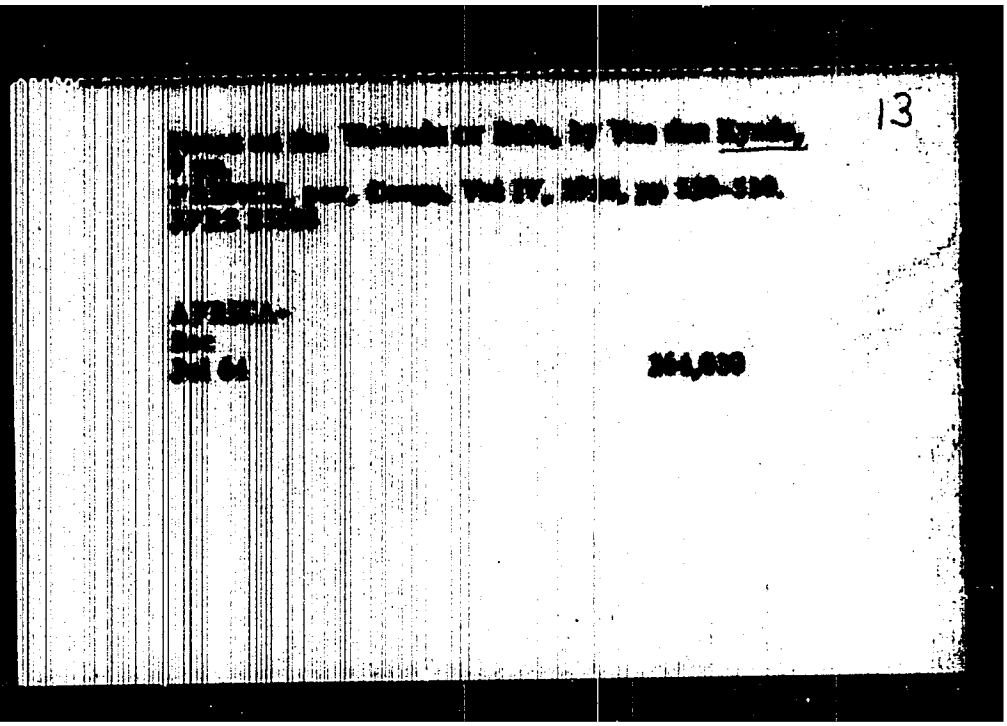

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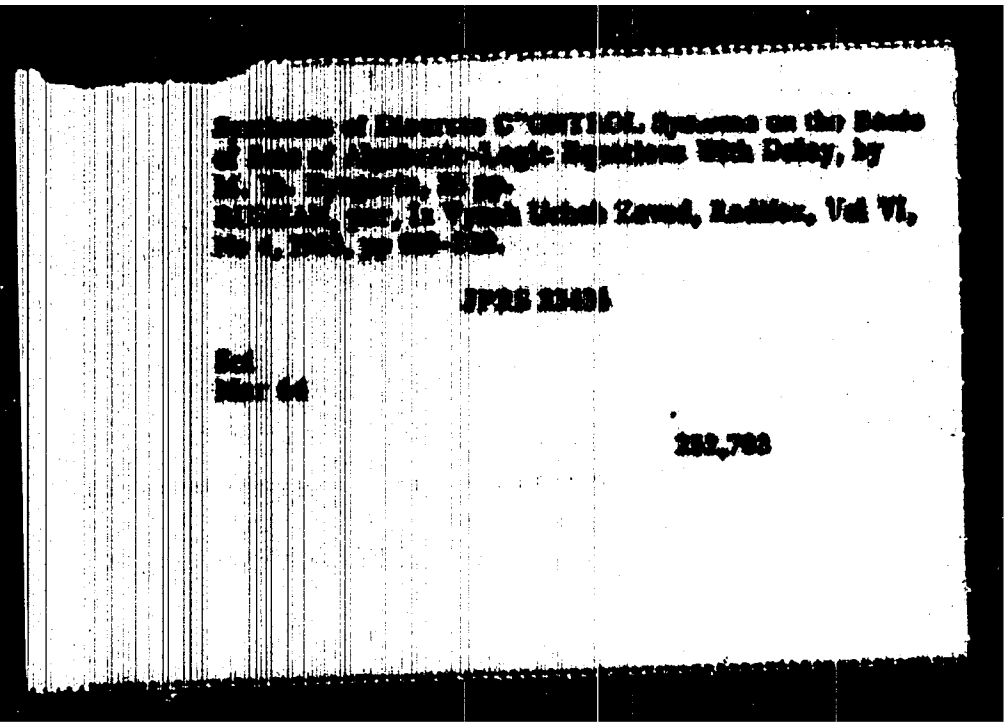

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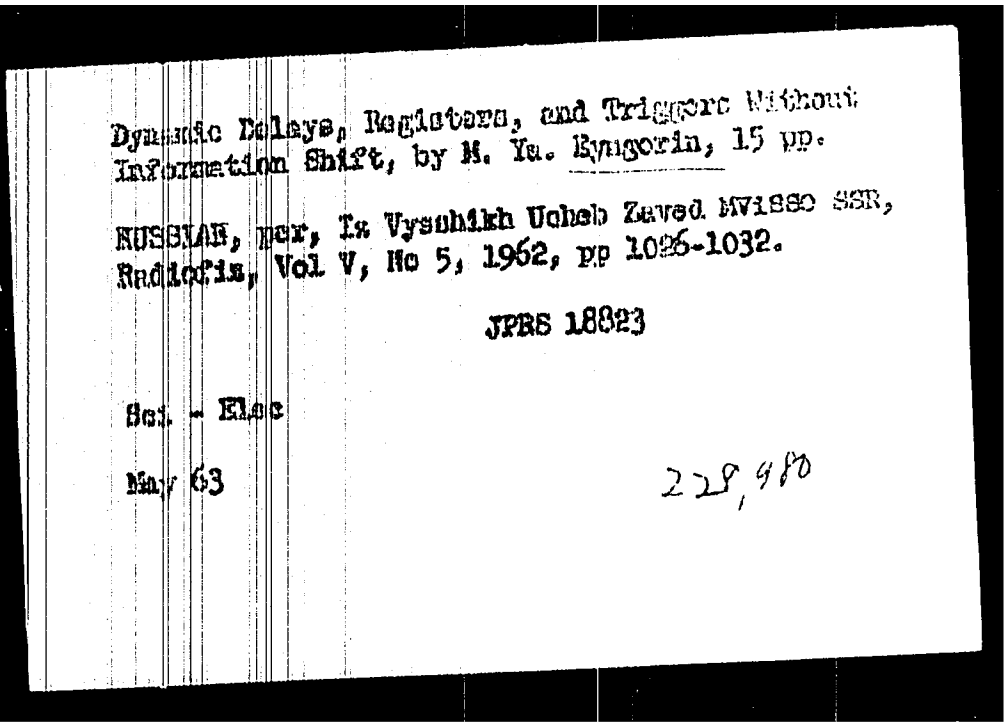

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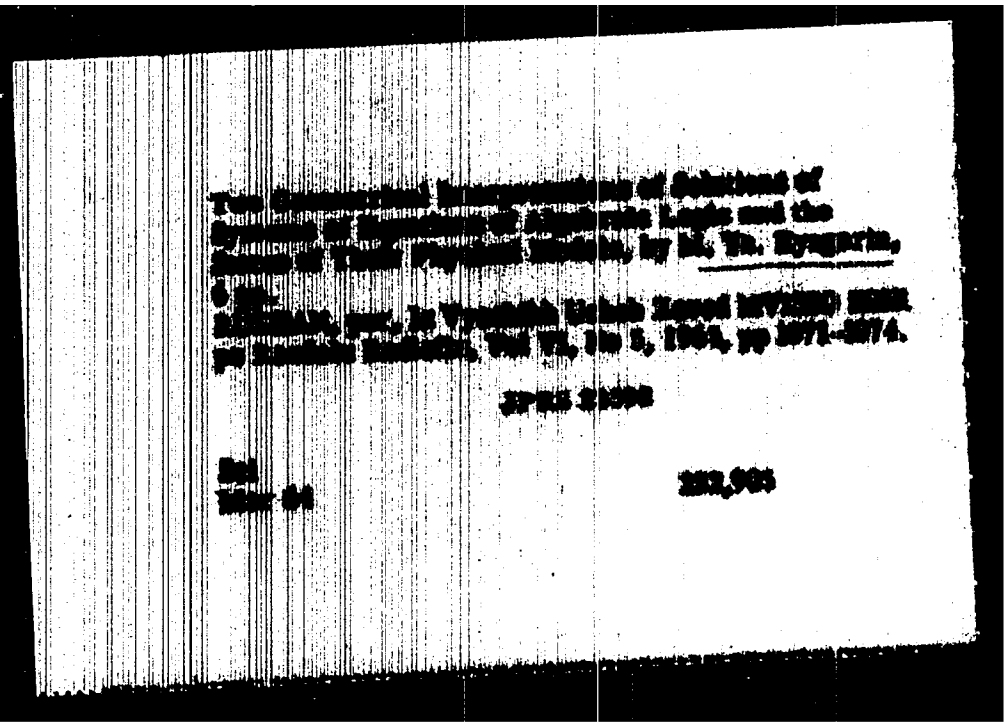

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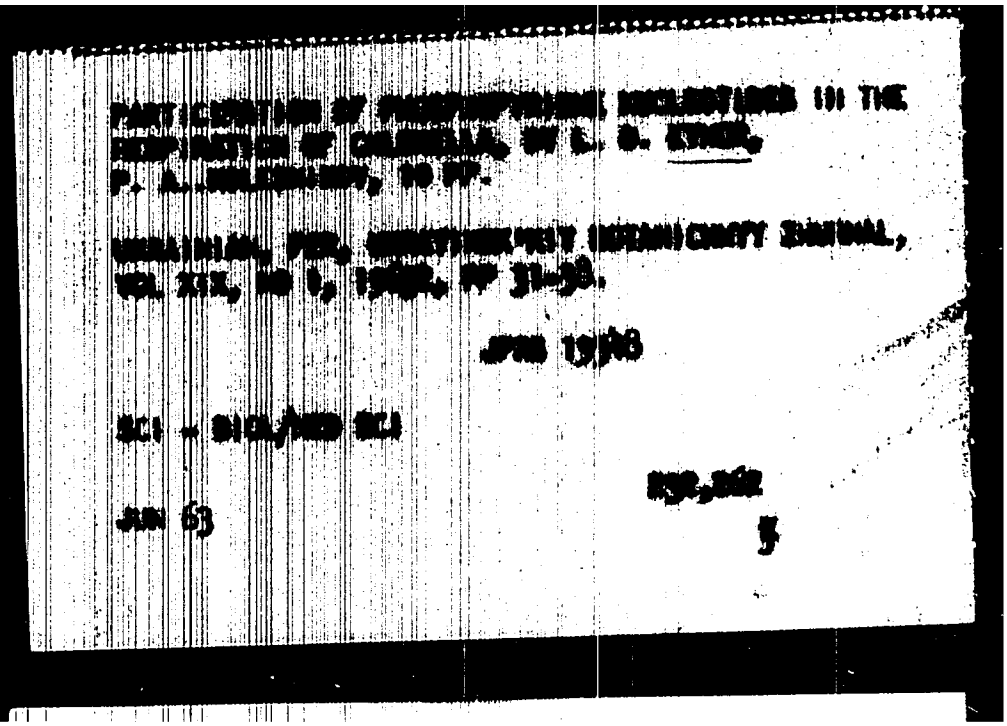

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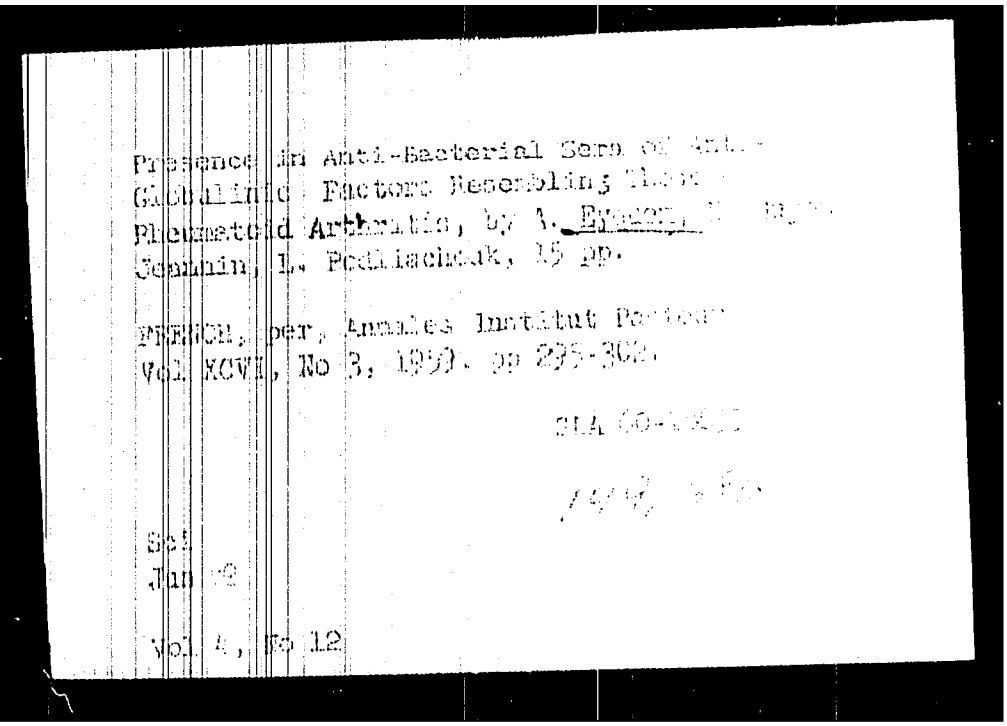

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Order from SLA mi$2.70, ph$4.80 61-10862

Trans. of [Société Chimique de France. Bulletin, 1960 [no. 2] p. 313-321.

The physicochemical study of sodium aluminate solutions of molecular caustic ratio at any level greater than one, by light diffusion, Raman spectrophotometry, and translation diffusion militates in favor of a unique aluminate ion, $Al(OH)_4^-$. A consequence of this fact is the possibility of a thermodynamic study of the Bayer phenomenon as represented by the reaction:
$2Al(OH)_3 + [Al(OH)_4]^- \rightleftharpoons (n + 1)\ Al(OH)_3 + OH^-$
Furthermore, it is shown that the separation of crystalline trihydrate cannot be spontaneous. This excludes the macromolecular or macro-ionic intermediate.
(Author)

(Chemistry--Physical, TT, v. 5, no. 12)

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1. Sodium aluminates--Molecular structure
2. Sodium aluminates--Reduction
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II. Lanaspeze, P.

Office of Technical Services

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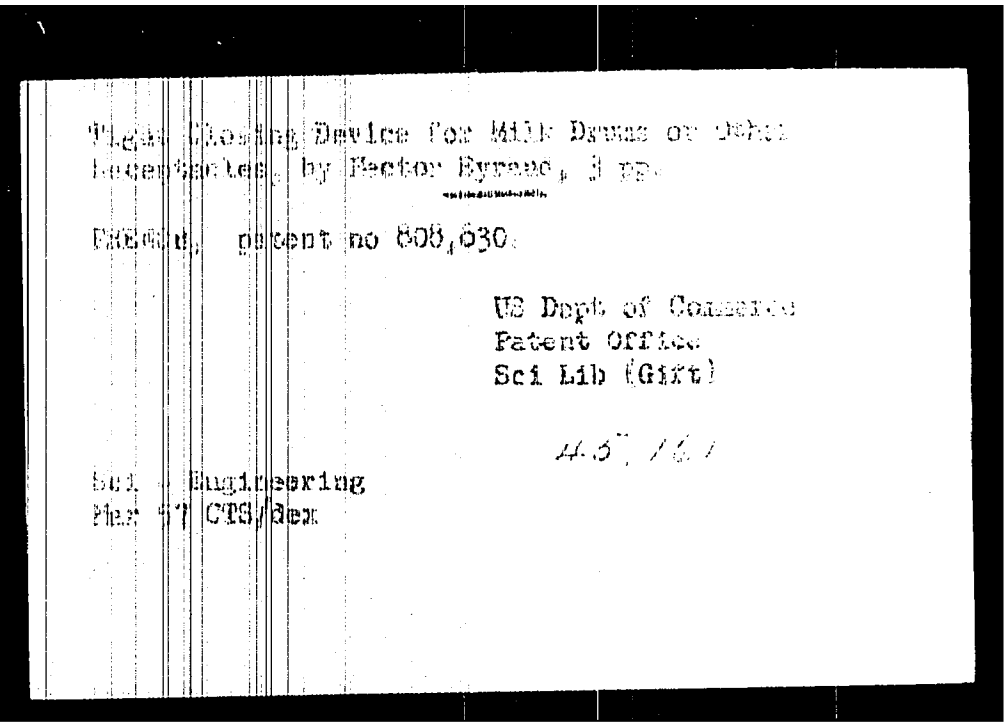

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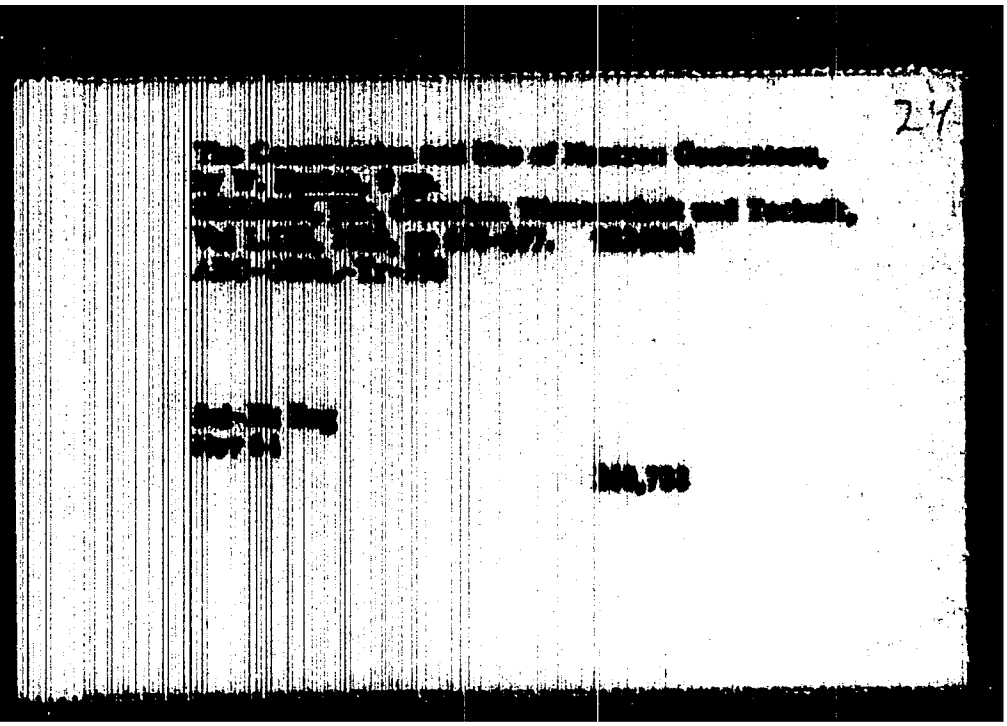

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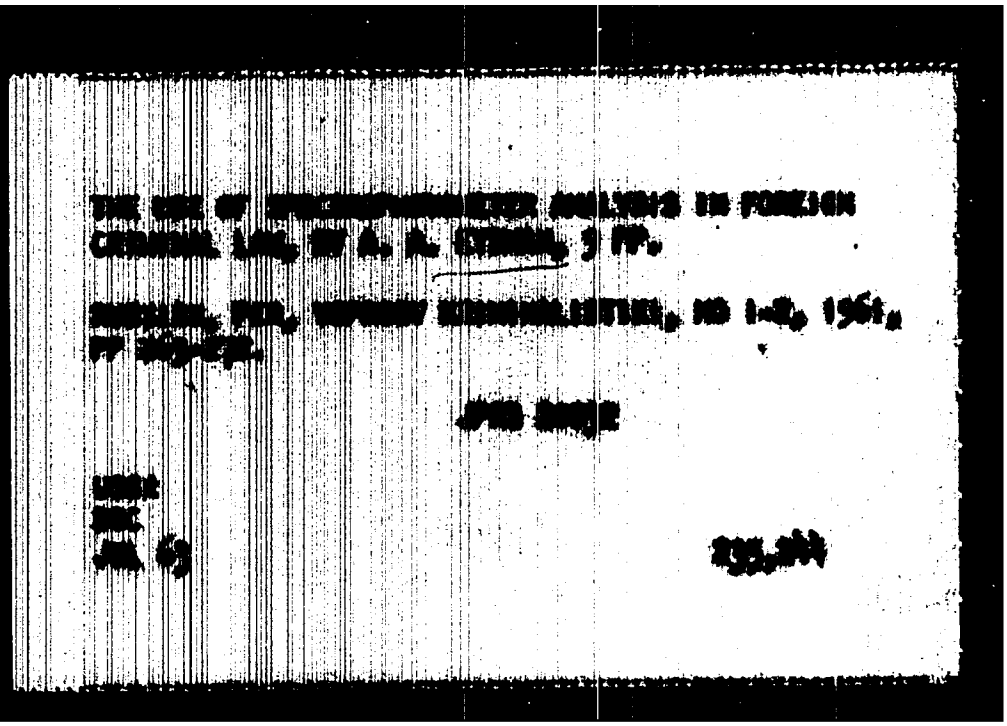

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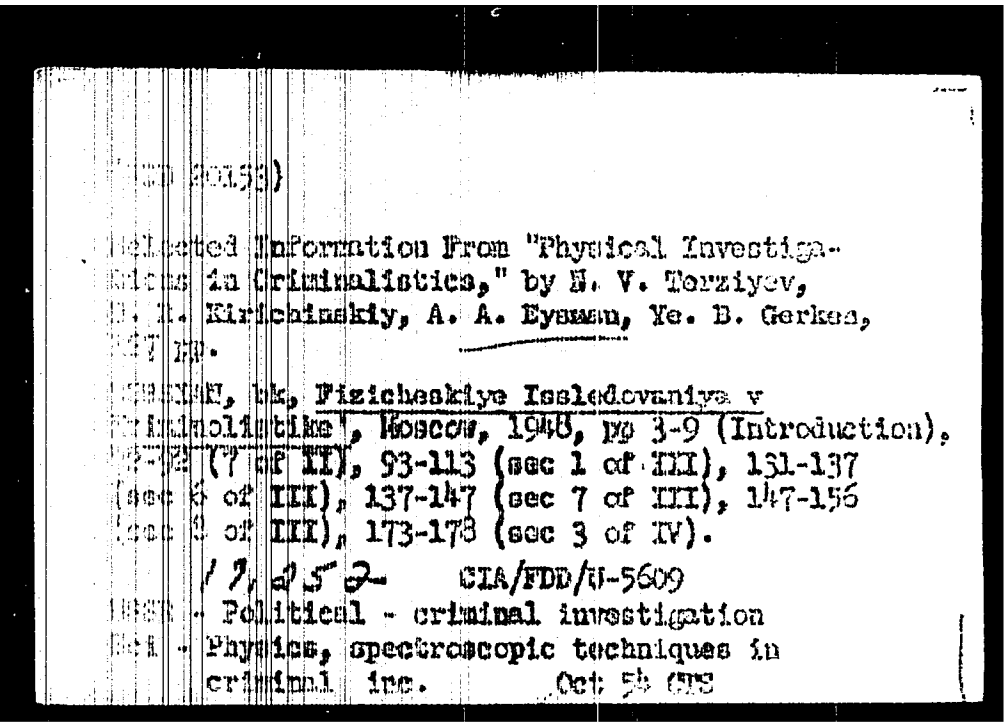

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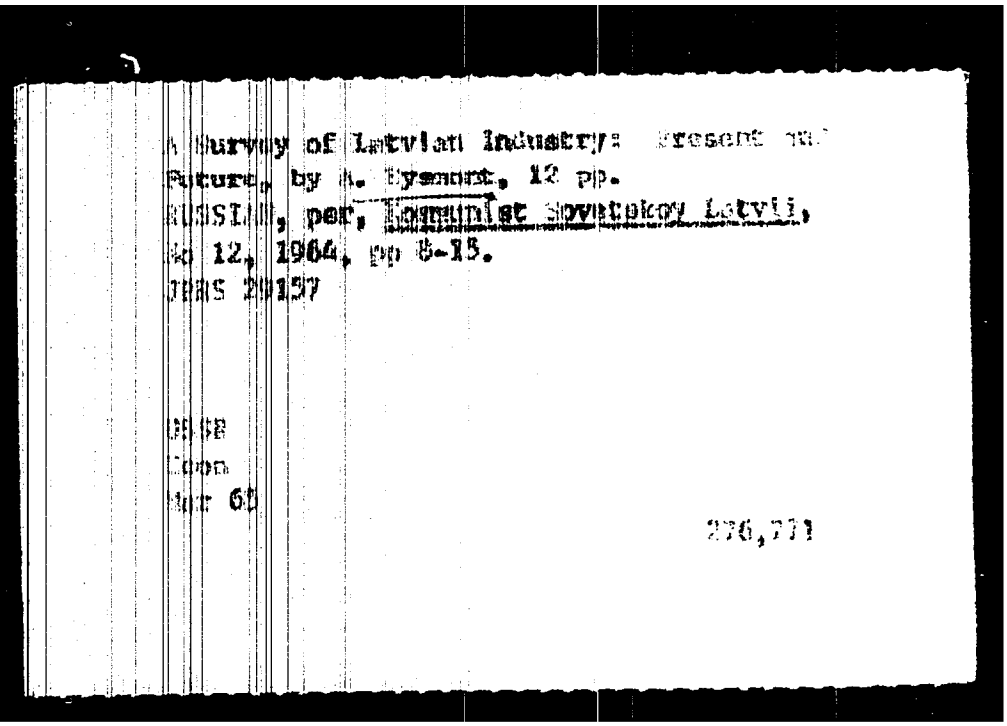

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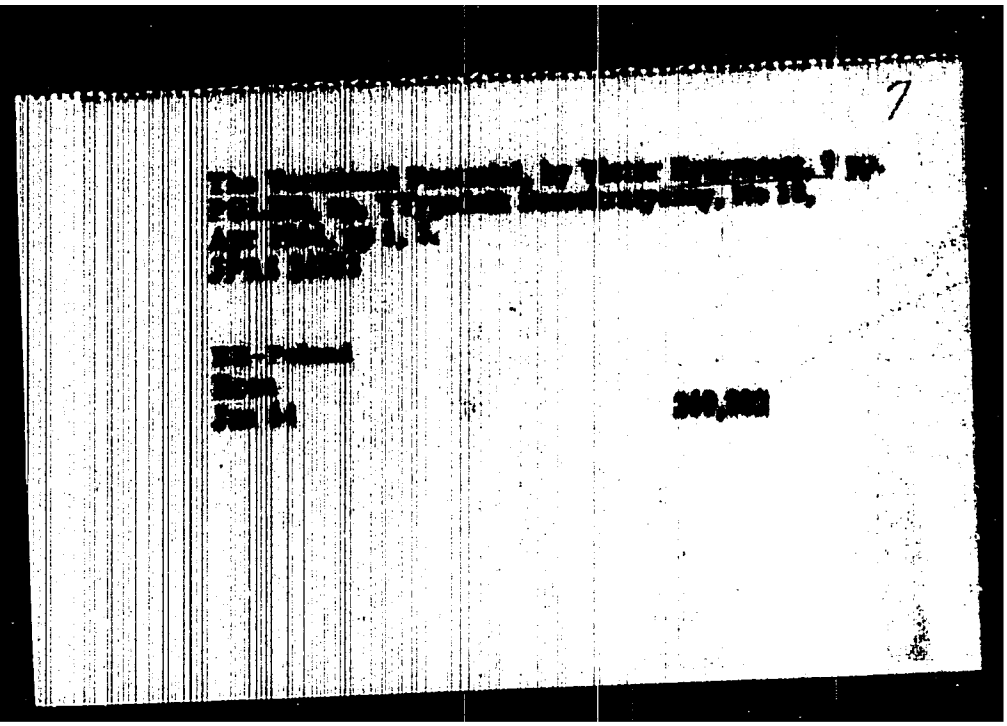

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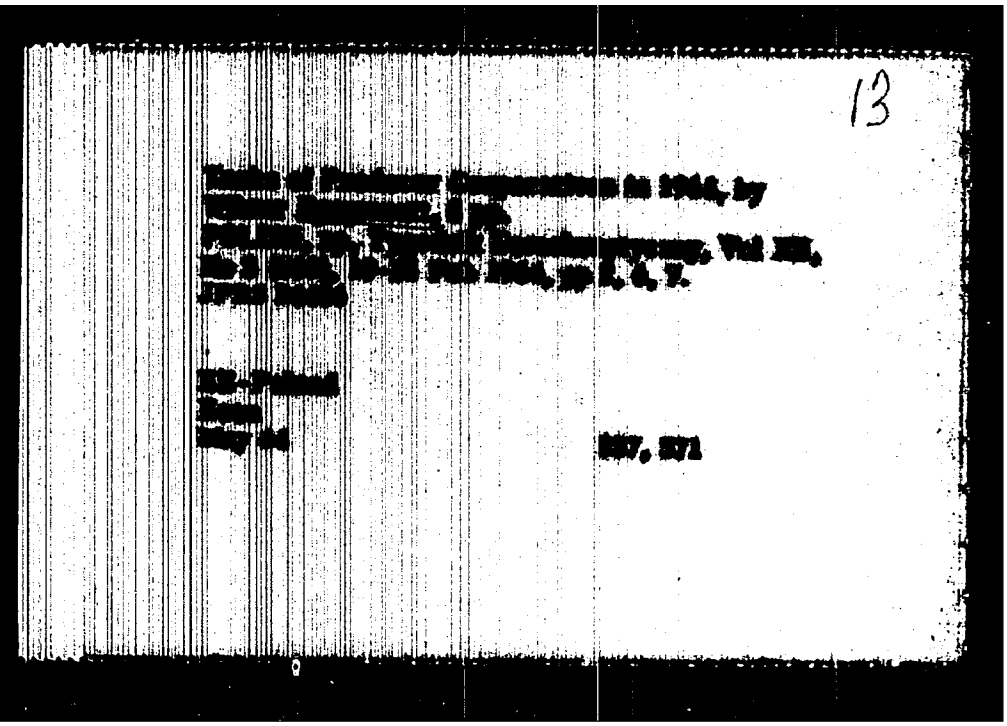

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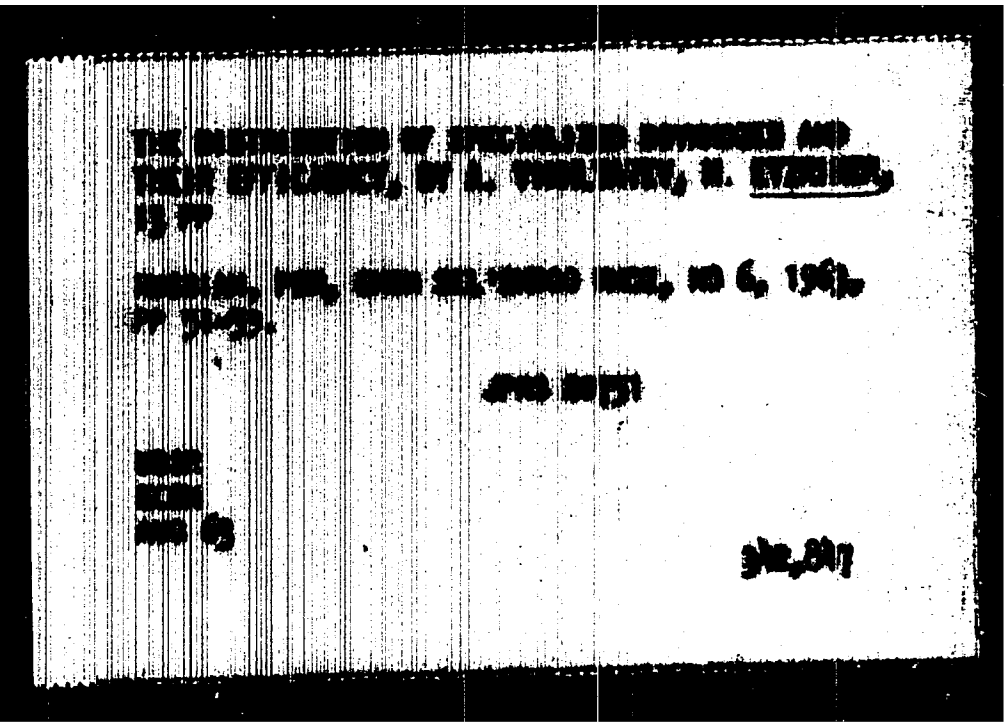

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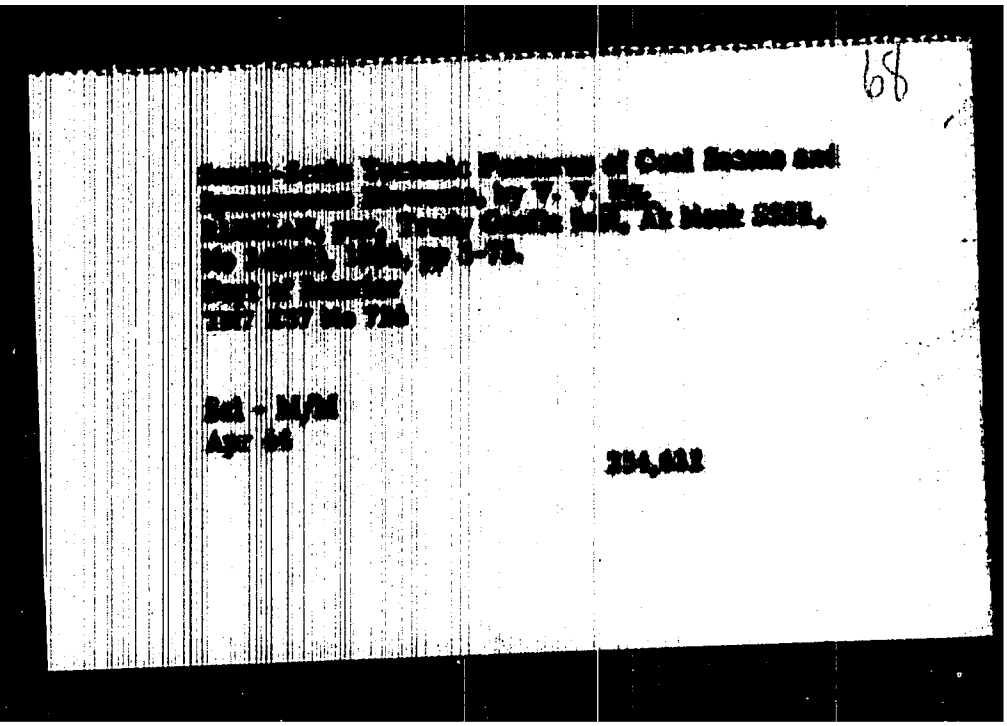

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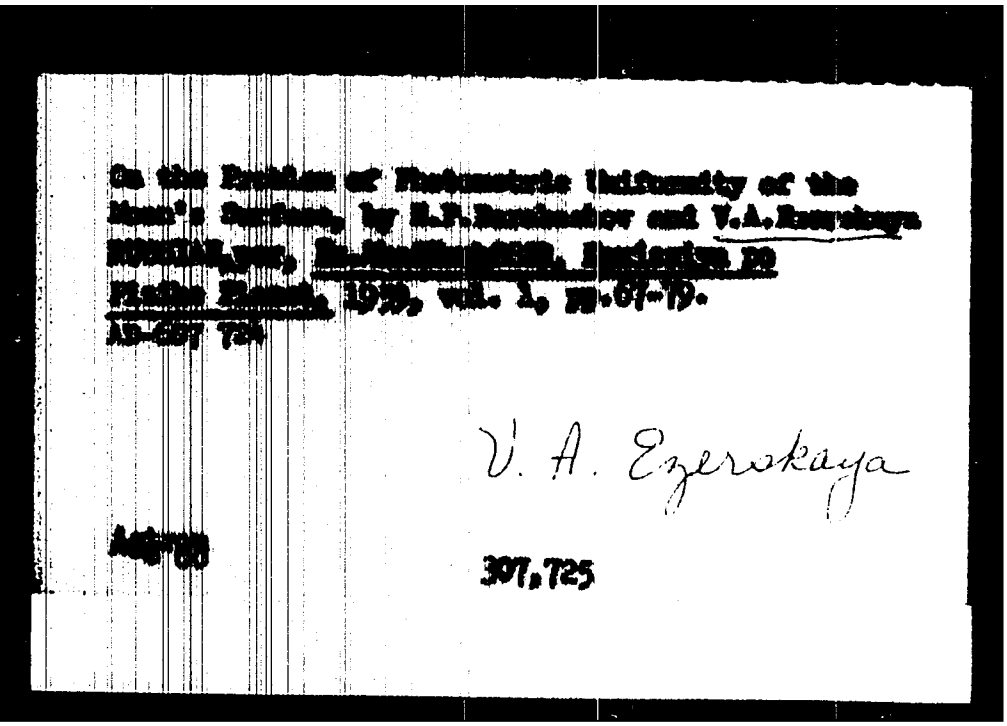

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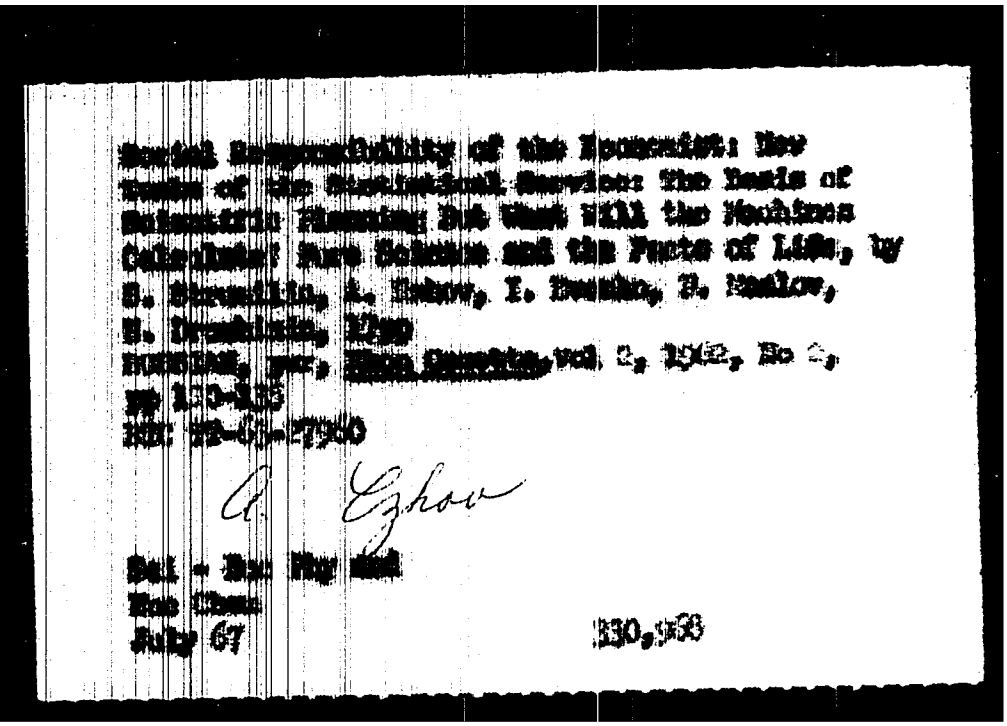

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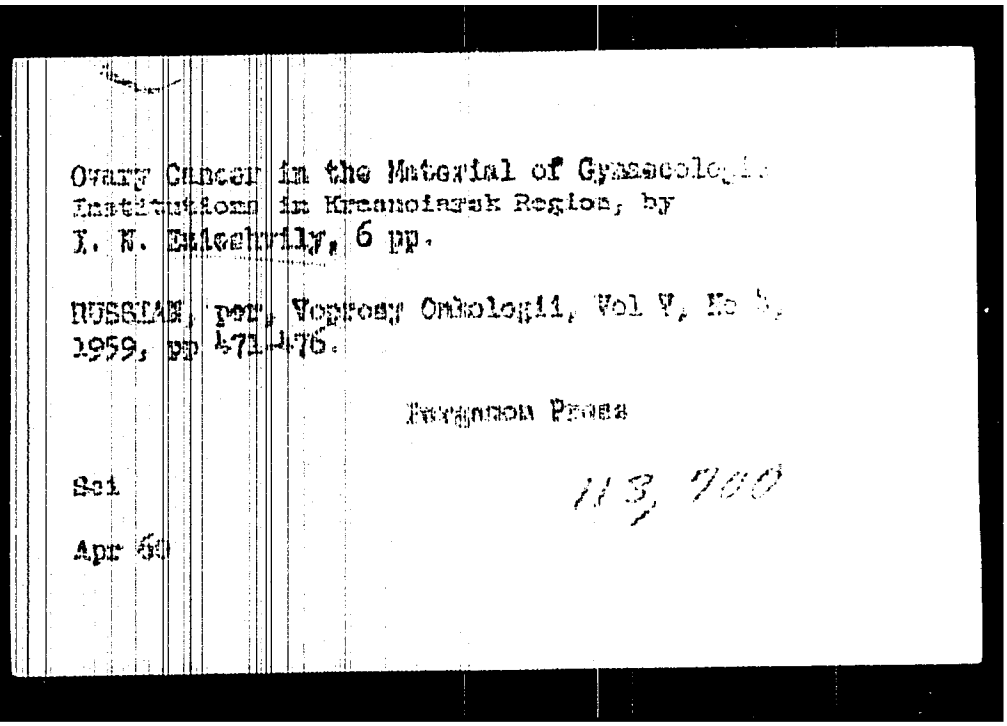

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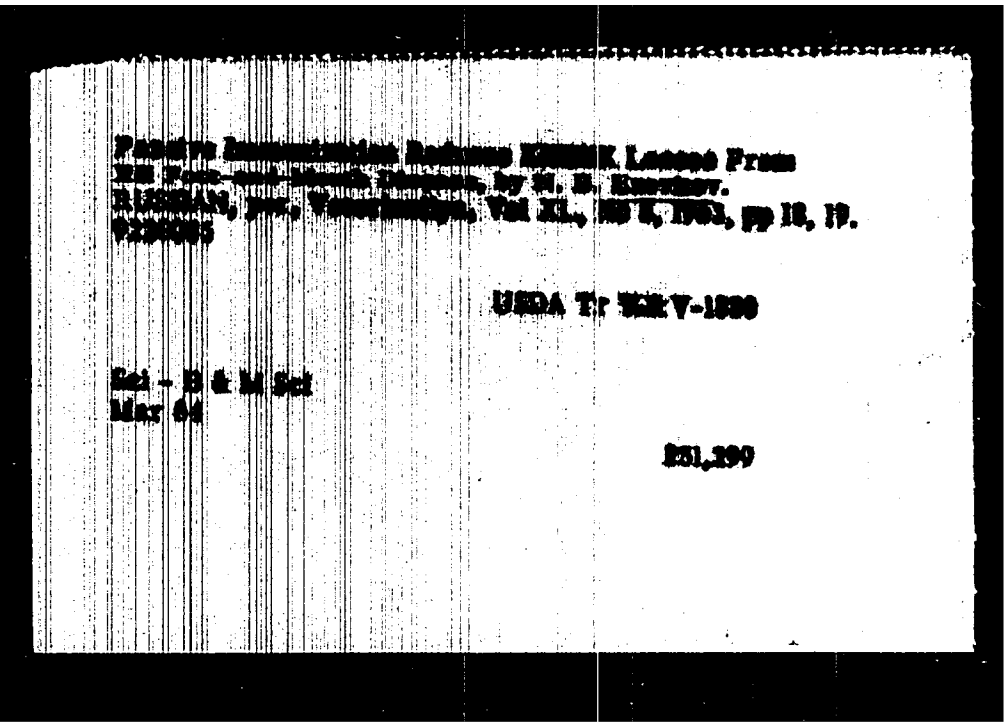

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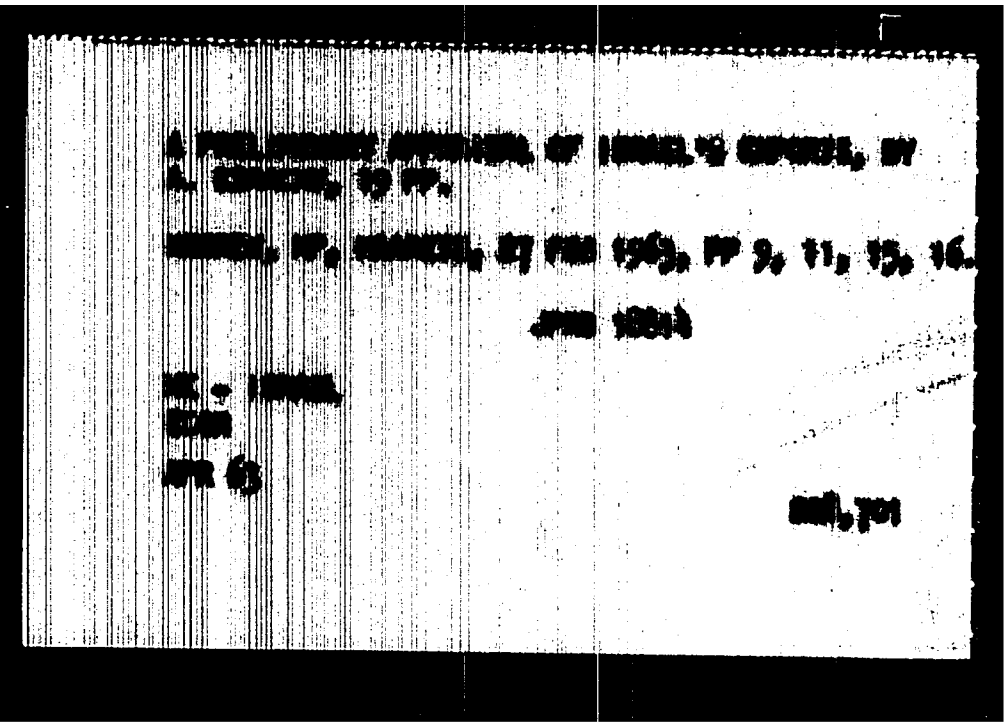

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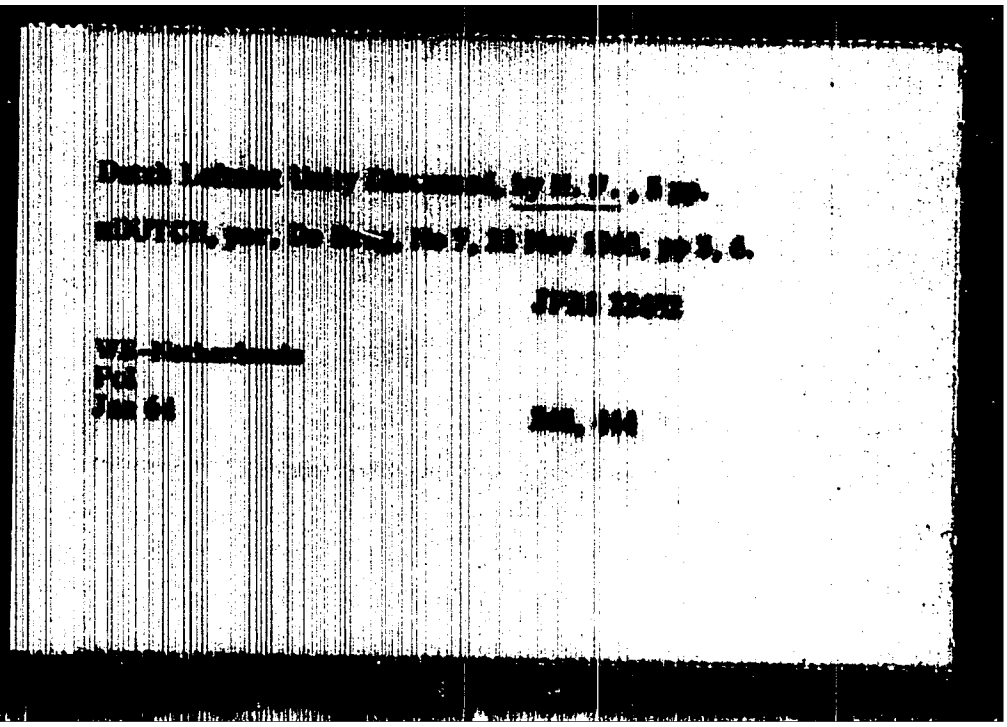

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